**Chemistry Unit 3 Atoms and the Periodic Table Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Bohr Model Worksheet**

Locate the following Elements on the Periodic Table. Find the **Atomic Number, Z.** Determine how many **Protons** and **Electrons** are in the Atom. Use the table of isotopes to determine the most abundant isotope **mass number, A**. Determine the number of **Neutrons** for the most abundant isotope. Once you have found your answers **draw a Bohr model** representing your Atom placing the electrons, protons, and neutrons where they go in the diagram.

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| 1. **Sodium** –  **Z = 11**  **Most abundant isotope A = 23**  **#p = 11 #e = 11 #n = 12**    M  L  e  K  e  e  e  e  e  e  e  p  n  e  e  e | 2. **Beryllium** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** |
| 3. **Sulfur** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** | 4. **Fluorine** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** |
| 5. **Calcium** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** | 6. **Argon** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** |

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| 7. **Lithium** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** | 8. **Nitrogen** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** |
| 9. **Aluminum** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** | 10. **Silicon** –  **Z =**  **Most abundant isotope A =**  **#p = #e = #n =** |